

# Activity #2 - Solution

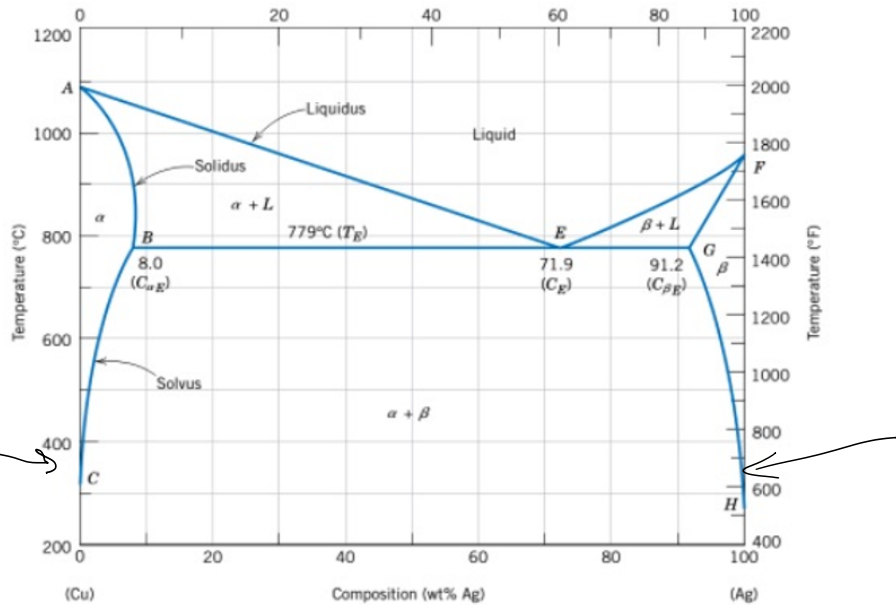
Thursday, September 3, 2020

10:14 AM

Activity #2 "Solution"

Our alloy is  
95% Cu & 5% Ag  
↳ both wt. %.

At room temp  
(I use 20°C), we  
have 2 phase  
system. Since  
the solvus curves  
reach the boundaries  
at points C & H, well  
above 20°C, then  
there is no solubility  
of Cu in Ag or Ag  
in Cu. This means  $\alpha$  is pure Cu &  $\beta$  is pure Ag  
so don't need to consider impurity scattering within  
 $\alpha$  or  $\beta$ .



$$\text{At } 20^\circ\text{C} \quad \rho_{\text{Cu}}^{20} = 15.7 \text{ n/cm} \left[ 1 + \frac{1}{232} (20 \text{ K}) \right] = 17.053 \text{ n/cm}$$

$$\rho_{\text{Ag}}^{20} = \dots 15.797 \text{ n/cm}$$

These are similar, so can use series mixtures rule

$$\rho_{\text{cell}}^{20} = \kappa_{\text{Cu}} \rho_{\text{Cu}}^{20} + \kappa_{\text{Ag}} \rho_{\text{Ag}}^{20}$$

⇒ Need volume fractions ↑  
Since  $\kappa_{\text{Cu}} + \kappa_{\text{Ag}} = 1$ , just find one:  $\kappa_{\text{Cu}} = 1 - \kappa_{\text{Ag}}$

$$\kappa_{\beta} = \frac{w_{\beta} \rho_{\text{alloy}}}{\rho_{\beta}} \Rightarrow \frac{w_{\text{Ag}} \rho_{\text{alloy}}}{\rho_{\text{Ag}}} \quad \text{since } \beta \text{ is pure Ag.}$$

↳ densities  $\rho$

$$\rho_{\text{alloy}} = \left[ \frac{w_{\text{Ag}}}{\rho_{\text{Ag}}} + \frac{w_{\text{Cu}}}{\rho_{\text{Cu}}} \right]^{-1} \Rightarrow \rho_{\text{alloy}} = 9.0262 \text{ g/cm}^3$$

$$\rho = 10.50 \text{ g/cm}^3 \quad \rho = 8.96 \text{ g/cm}^3$$

$$\Rightarrow X_{\text{Ag}} = 0.04298$$

$$\rho_{\text{pell}}^{20} = 16.999 \text{ n}\Omega\text{m}$$

← almost the same as pure Cu, which might be expected.

Can use Reynolds & Hough to check series mixtures rule.  
 ↳ get 16.996 nΩm, so mixtures rule is pretty good.

At 800°C, have a single phase alloy isomorphous solid solution. So now Ag is dissolved in Cu & acts as impurity.  $\Rightarrow \rho_{\text{pell}}^{800} = \rho_{\text{Cu}}^{800} + C X_{\text{Ag}} (1 - X_{\text{Ag}})$  ← atomic fraction.

$$\rho_{\text{Cu}}^{800} = 69.838 \text{ n}\Omega\text{m}$$

$$X_{\text{Ag}} = \frac{w_{\text{Ag}}/m_{\text{Ag}}}{\frac{w_{\text{Ag}}}{m_{\text{Ag}}} + \frac{w_{\text{Cu}}}{m_{\text{Cu}}}} = 0.03007$$

$\uparrow$  107.87 g/mol       $\uparrow$  63.55 g/mol

With  $C = 2800 \text{ n}\Omega\text{m}$  for Nordheim coeff.

$$\Rightarrow \rho_{\text{pell}}^{800} = 152 \text{ n}\Omega\text{m} \quad \leftarrow \text{much higher than } 20^\circ\text{C}.$$

So what's the change in power?

$$\% \Delta \text{ power} = 100 \left( \frac{P_{800} - P_{20}}{P_{20}} \right) = \frac{I^2 R_{800} - I^2 R_{20}}{I^2 R_{20}}$$

↳ I assume the current stays constant.

$$= \% \Delta \text{ power} = \left( \frac{\rho_{\text{pell}}^{800} - \rho_{\text{pell}}^{20}}{\rho_{\text{pell}}^{20}} \right) 100 = \boxed{791\%} !!$$

That's too much of a change.